

Constantinos D. Antoniadis,^a
Alexander J. Blake,^{b*} Sotiris K.
Hadjikakou,^a Nick Hadjiliadis,^a
Peter Hubberstey,^b Martin
Schröder^b and Claire Wilson^b

^aSection of Inorganic and Analytical Chemistry,
Department of Chemistry, University of
Ioannina, 45110 Ioannina, Greece, and ^bSchool
of Chemistry, The University of Nottingham,
University Park, Nottingham NG7 2RD, England

Correspondence e-mail:
a.j.blake@nottingham.ac.uk

Structural characterization of selenium and selenium-diiodine analogues of the antithyroid drug 6-*n*-propyl-2-thiouracil and its alkyl derivatives

Received 8 March 2006
Accepted 29 March 2006

The structures of four selenium analogues of the antithyroid drug 6-*n*-propyl-2-thiouracil [systematic name: 2,3-dihydro-6-*n*-propyl-2-thioxopyrimidin-4(1*H*)-one], namely 6-methyl-2-selenouracil, C₅H₆N₂OSe (1), 6-ethyl-2-selenouracil, C₆H₈N₂OSe (2), 6-*n*-propyl-2-selenouracil, C₇H₁₀N₂OSe (3), and 6-isopropyl-2-selenouracil, C₇H₁₀N₂OSe (4), are described, along with that of the dichloromethane monosolvate of 6-isopropyl-2-selenouracil, C₇H₁₀N₂OSe·CH₂Cl₂ (4·CH₂Cl₂). The extended structure of (1) is a two-dimensional sheet of topology 6³ with a brick-wall architecture. The extended structures of (2) and (4) are analogous, being based on a chain of eight-membered $R_8^6(32)$ hydrogen-bonded rings. In (3) and (4·CH₂Cl₂), $R_2^2(8)$ hydrogen bonding links molecules into chains. 6-*n*-Propyl-2-selenouracil·I₂, C₇H₁₀N₂OSe·I₂ (7), is a charge-transfer complex with a 'spoke' structure, the extended structure of which is based on a linear chain formed principally by intermolecular N—H···O hydrogen bonds. Re-crystallization of 6-ethyl-2-selenouracil or (7) from acetone gave crystals of the diselenides [N-(6'-ethyl-4'-pyrimidone)(6-ethyl-2-selenouracil)₂(Se—Se)]·2H₂O (9·2H₂O) or [N-(6'-*n*-propyl-4'-pyrimidone)(6-*n*-propyl-2-selenouracil)₂(Se—Se)] (10), respectively: these have similar extended chain structures formed *via* N—H···O and C—H···O hydrogen bonds, stacked to give two-dimensional sheets. Re-crystallization of (7) from methanol/ acetonitrile led *via* deselenation to the formation of crystals of 6-*n*-propyl-2-uracil (11), in which six symmetry-related molecules combine to form a six-membered $R_6^6(24)$ hydrogen-bonded ring, with each pair of molecules linked by an $R_2^2(8)$ motif.

1. Introduction

Prior to 1960, no uncharged covalent compounds between selenium and iodine had been discovered (Dasent, 1965). It was during the decade 1960–1969 that the structures of the first selenoether–iodine complexes were first reported (Chao & McCullough, 1961; Hope & McCullough, 1962; Maddox & McCullough, 1966; Bjorvatten, 1963; Dahl & Hassel, 1965; Holmesland & Römning, 1966; Bent, 1968). Currently, organic selones and/or selenoamides are recognized as potential donors towards diiodine and some iodine-containing compounds, generating charge-transfer complexes that are generally more stable than those of the corresponding sulfur ligands (*e.g.* du Mont *et al.*, 2001). The reactions between selones or selenoamides with diiodine I₂ or interhalogens I—X (X = Br or Cl) lead to the formation of iodine charge-transfer complexes (Aragoni *et al.*, 1999, and references therein), which adopt a number of structures including:

(i) 'spoke structures' or 'extended spoke structures' bearing a linear arrangement of Se—I—X ($X = \text{I}, \text{Br}$ or Cl) or Se—I—I ··· I—I groups;

(ii) two-coordinate iodine(I) cationic or iodonium salts with seleno ligands ($[\text{LSe}-\text{X}-\text{SeL}]$);

(iii) donor oxidation products including dicationic diselenides $[\text{LSe}-\text{SeL}]^{2+} \cdot 2\text{I}_3^-$ or neutral diselenides; and

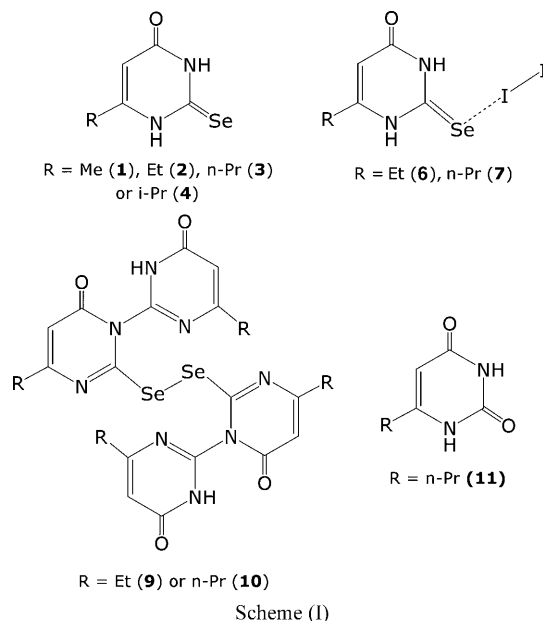
(iv) T-shaped compounds containing the linear group I—Se—X ($X = \text{I}, \text{Br}$ or Cl ; Aragoni *et al.*, 2001, and references therein).

The mono-cationic diselenide geometry is the only interaction not yet observed in products of reactions between iodine and selenes. As part of a study of charge-transfer complexes with iodine and polyiodides (Blake *et al.*, 1995; Blake, Devillanova, Garau *et al.*, 1998; Blake, Devillanova, Gould *et al.*, 1998; Blake, Li *et al.*, 1998), we report herein the structures of 6-*n*-propyl-2-selenouracil- I_2 , 6-*n*-propyl-2-selenouracil itself, some selenium analogues of 6-alkyl-2-thiouracils, two diselenides and the deselenation product 6-*n*-propyl-2-uracil.

2. Experimental

2.1. Synthesis and crystal growth

The preparative details for all nine compounds [see Scheme (I)] have been reported previously (Antoniadis *et al.*, 2006). Compounds (1), (2), (3) and (4) were re-crystallized from water to give colourless crystals, while re-crystallization of (4)



from dichloromethane gave crystals of (4·CH₂Cl₂). Crystals of (7) were grown from chloroform solutions, while re-crystallization of (6) and (7) from acetone gave the oxidation products (9·2H₂O) and (10), respectively. Re-crystallization of (7) from methanol/acetonitrile led *via* deselenation to crystals of 6-*n*-propyl-2-uracil (11).

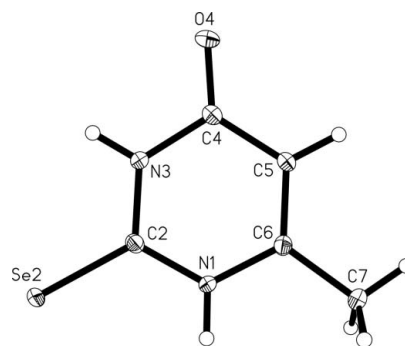


Figure 1

Displacement ellipsoid plot of (1) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level.

2.2. Data collection, structure solution and refinement

Details of cell data, data collection and structure solution and refinement are summarized in Table 1¹. Except for (9·2H₂O), which was solved using *SIR92* direct methods (Altomare *et al.*, 1994), the structures were solved by direct methods using *SHELXS97* (Sheldrick, 1990) and developed by difference Fourier methods using *SHELXL97* (Sheldrick, 1998). With the exception of the *sp*²-bound methyl group in (1), where they were located from a circular difference Fourier synthesis and refined as part of a rigid rotating group, all carbon- and nitrogen-bound H atoms were placed in geometrically calculated positions and refined using riding models (*SHELXL97*; Sheldrick, 1998). Crystals of (9·2H₂O) are affected by twinning, principally by a 180° rotation about [001] for which the twin fraction was 0.291 (2). The water H atoms in (9·2H₂O) were not located.

The structure determination of (3) posed particular difficulties: the crystals were very small in two dimensions (10 and 20 μm) and efforts to obtain a dataset using a sealed X-ray tube or a rotating-anode source gave no significant diffraction. However, we were able to obtain a weak dataset on Station 9.8 of the Daresbury Synchrotron Radiation Source, although even this did not overcome all of the limitations imposed by the poor crystal quality. We found that only the Se atoms could be refined with anisotropic displacement parameters, and a total of 99 geometric, planarity and similarity restraints were necessary for a stable refinement. Although we cannot therefore discuss the fine details of the molecular geometry, it did prove possible to obtain reliable information about the extended structure of (3).

3. Results and discussion

3.1. Molecular structures of (1)–(4), (4·CH₂Cl₂) and (11)

Displacement ellipsoid plots of (1)–(4), (4·CH₂Cl₂) and (11) are shown in Figs. 1–6 while selected interatomic distances and angles are listed in Tables 2 and 4. Although the asymmetric

¹ Supplementary data for this paper are available from the IUCr electronic archives (Reference: SO5003). Services for accessing these data are described at the back of the journal.

Table 1
Experimental table.

	(1)	(2)	(3)	(4)	(4-CH ₂ Cl ₂)
Crystal data					
Chemical formula	C ₅ H ₆ N ₂ OSe	C ₆ H ₈ N ₂ OSe	C ₇ H ₁₀ N ₂ OSe	C ₇ H ₁₀ N ₂ OSe	C ₈ H ₁₂ Cl ₂ N ₂ OSe
<i>M_r</i>	189.08	203.10	217.13	217.13	302.06
Cell setting, space group	Monoclinic, <i>P</i> 2 ₁ / <i>c</i>	Triclinic, <i>P</i> $\bar{1}$	Orthorhombic, <i>Pbca</i>	Triclinic, <i>P</i> $\bar{1}$	Triclinic, <i>P</i> $\bar{1}$
Temperature (K)	120 (2)	150 (2)	120 (2)	150 (2)	150 (2)
<i>a</i> , <i>b</i> , <i>c</i> (Å)	4.3411 (7), 14.756 (2), 9.690 (2)	8.394 (2), 10.029 (2), 14.931 (4)	10.568 (7), 11.257 (7), 28.79 (2)	8.9192 (9), 10.6403 (10), 15.1965 (15)	8.841 (3), 11.259 (3), 12.424 (4)
α , β , γ (°)	90.00, 90.157 (2), 90.00	101.023 (4), 100.893 (4), 105.705 (4)	90.00, 90.00, 90.00	106.019 (2), 105.366 (2), 96.166 (2)	90.450 (5), 105.350 (4), 92.945 (5)
<i>V</i> (Å ³)	620.71 (18)	1148.5 (5)	3425 (4)	1311.0 (2)	1190.7 (6)
<i>Z</i>	4	6	16	6	4
<i>D_x</i> (Mg m ⁻³)	2.023	1.762	1.684	1.650	1.685
Radiation type	Synchrotron	Mo <i>K</i> α	Synchrotron	Mo <i>K</i> α	Mo <i>K</i> α
No. of reflections for cell parameters	4328	2232	3142	3689	2249
θ range (°)	2.6–29.0	2.6–24.8	2.6–27.5	2.4–27.6	2.4–24.6
μ (mm ⁻¹)	5.96	4.84	4.33	4.24	3.57
Crystal form, colour	Needle, colourless	Triangular prism, colourless	Lath, colourless	Tablet, colourless	Plate, colourless
Crystal size (mm)	0.10 × 0.01 × 0.01	0.21 × 0.12 × 0.04	0.20 × 0.02 × 0.01	0.25 × 0.14 × 0.06	0.36 × 0.22 × 0.02
Data collection					
Diffractometer	Bruker SMART APEXII CCD diffractometer	Bruker SMART APEX CCD area detector	Bruker SMART APEXII CCD diffractometer	Bruker SMART APEX CCD area detector	Bruker SMART APEX CCD area detector
Data collection method	Fine-slice ω scans	ω scans	Fine-slice ω scans	ω scans	ω scans
Absorption correction	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)
<i>T_{min}</i>	0.816	0.687	0.260	0.581	0.684
<i>T_{max}</i>	1.000	1.000	1.000	0.770	1.000
No. of measured, independent and observed reflections	6350, 1770, 1642	8174, 4025, 3189	24 855, 3471, 2417	8163, 5780, 4932	10 743, 6193, 4391
Criterion for observed reflections	<i>I</i> > 2 σ (<i>I</i>)	<i>I</i> > 2 σ (<i>I</i>)	<i>I</i> > 2 σ (<i>I</i>)	<i>I</i> > 2 σ (<i>I</i>)	<i>I</i> > 2 σ (<i>I</i>)
<i>R_{int}</i>	0.027	0.038	0.250	0.013	0.050
θ_{\max} (°)	29.0	25.1	25.0	27.6	27.5
Range of <i>h</i> , <i>k</i> , <i>l</i>	−6 ⇒ <i>h</i> ⇒ 6 −20 ⇒ <i>k</i> ⇒ 21 −13 ⇒ <i>l</i> ⇒ 13	−10 ⇒ <i>h</i> ⇒ 10 −11 ⇒ <i>k</i> ⇒ 11 −17 ⇒ <i>l</i> ⇒ 17	−13 ⇒ <i>h</i> ⇒ 13 −14 ⇒ <i>k</i> ⇒ 14 −35 ⇒ <i>l</i> ⇒ 35	−11 ⇒ <i>h</i> ⇒ 7 −13 ⇒ <i>k</i> ⇒ 13 −17 ⇒ <i>l</i> ⇒ 19	−11 ⇒ <i>h</i> ⇒ 11 −14 ⇒ <i>k</i> ⇒ 14 −16 ⇒ <i>l</i> ⇒ 16
Refinement					
Refinement on	<i>F</i> ²	<i>F</i> ²	<i>F</i> ²	<i>F</i> ²	<i>F</i> ²
<i>R</i> [<i>F</i> ² > 2 σ (<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.025, 0.074, 0.80	0.056, 0.146, 1.03	0.173, 0.408, 1.18	0.028, 0.071, 1.04	0.052, 0.099, 0.97
No. of reflections	1770	4025	3471	5780	6193
No. of parameters	83	271	99	298	254
H-atom treatment	Rigid rotating group; riding model	Constrained to parent site	Constrained to parent site	Constrained to parent site	Constrained to parent site
Weighting scheme	$w = 1/[\sigma^2(F_o^2) + (0.079P)^2]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.094P)^2]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.025P)^2 + 367.0]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.033P)^2 + 0.854P]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.0305P)^2]$, where $P = (F_o^2 + 2F_c^2)/3$
(Δ/σ) _{max}	0.001	0.001	0.002	0.001	0.001
$\Delta\rho_{\max}$, $\Delta\rho_{\min}$ (e Å ⁻³)	0.57, −0.80	1.48, −1.02	2.38, −2.33	0.65, −0.40	0.87, −0.51
	(7)	(9)	(10)	(11)	
Crystal data					
Chemical formula	C ₇ H ₁₀ I ₂ N ₂ OSe	C ₂₄ H ₂₆ N ₈ O ₄ Se ₂ ·2H ₂ O	C ₂₈ H ₃₄ N ₈ O ₄ Se ₂	C ₇ H ₁₀ N ₂ O ₂	
<i>M_r</i>	470.93	684.48	704.55	154.17	
Cell setting, space group	Triclinic, <i>P</i> $\bar{1}$	Triclinic, <i>P</i> $\bar{1}$	Triclinic, <i>P</i> $\bar{1}$	Rhombohedral, <i>R</i> $\bar{3}$	
Temperature (K)	150 (2)	120 (2)	120 (2)	150 (2)	

Table 1 (continued)

	(7)	(9)	(10)	(11)
<i>a</i> , <i>b</i> , <i>c</i> (Å)	6.603 (2), 7.700 (3), 13.037 (5)	4.8330 (2), 9.7970 (5), 14.1796 (8)	5.0717 (6), 11.8615 (14), 11.9385 (14)	19.9444 (10), 19.9444 (10), 9.8918 (10)
α , β , γ (°)	75.969 (6), 86.808 (6), 73.113 (6)	83.490 (3), 84.431 (3), 89.353 (3)	83.161 (2), 82.785 (2), 84.358 (2)	90.00, 90.00, 120.00
<i>V</i> (Å ³)	615.3 (7)	663.91 (6)	704.90 (14)	3407.6 (10)
<i>Z</i>	2	1	1	18
<i>D_x</i> (Mg m ⁻³)	2.542	1.712	1.660	1.352
Radiation type	Mo <i>K</i> α	Mo <i>K</i> α	Synchrotron	Mo <i>K</i> α
No. of reflections for cell parameters	1665	2951	1672	1748
θ range (°)	2.8–27.0	2.9–27.5	2.2–27.0	2.4–27.4
μ (mm ⁻¹)	8.04	2.84	2.67	0.10
Crystal form, colour	Column, red	Needle, orange	Plate, colourless	Lens, orange
Crystal size (mm)	0.22 × 0.10 × 0.06	0.20 × 0.04 × 0.03	0.08 × 0.05 × 0.01	0.52 × 0.17 × 0.17
Data collection				
Diffraction	Bruker SMART1000 CCD area detector	Bruker Nonius kappaCCD area detector	Bruker SMART APEXII CCD diffractometer	Bruker SMART APEX CCD area detector
Data collection method	ω scans	ω and φ	Fine-slice ω scans	ω scans
Absorption correction	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)	Multi-scan (based on symmetry-related measurements)	None
<i>T</i> _{min}	0.114	0.679	0.804	–
<i>T</i> _{max}	0.209	1.000	1.000	–
No. of measured, independent and observed reflections	5463, 2660, 1702	13 298, 13 281, 12 058	8292, 4198, 3144	5827, 1750, 1297
Criterion for observed reflections	$I > 2\sigma(I)$	$I > 2\sigma(I)$	$I > 2\sigma(I)$	$I > 2\sigma(I)$
<i>R</i> _{int}	0.061	0.060	0.038	0.039
θ_{\max} (°)	27.5	27.7	28.9	27.5
Range of <i>h</i> , <i>k</i> , <i>l</i>	–8 ⇒ <i>h</i> ⇒ 8 –9 ⇒ <i>k</i> ⇒ 9 –16 ⇒ <i>l</i> ⇒ 16	–6 ⇒ <i>h</i> ⇒ 6 –12 ⇒ <i>k</i> ⇒ 12 –18 ⇒ <i>l</i> ⇒ 18	–7 ⇒ <i>h</i> ⇒ 7 –16 ⇒ <i>k</i> ⇒ 16 –17 ⇒ <i>l</i> ⇒ 17	–25 ⇒ <i>h</i> ⇒ 19 –24 ⇒ <i>k</i> ⇒ 25 –11 ⇒ <i>l</i> ⇒ 12
Refinement				
Refinement on	<i>F</i> ²	<i>F</i> ²	<i>F</i> ²	<i>F</i> ²
$R[F^2 > 2\sigma(F^2)]$, <i>wR</i> (<i>F</i> ²), <i>S</i>	0.038, 0.091, 0.87	0.084, 0.221, 1.04	0.042, 0.097, 0.98	0.044, 0.131, 1.05
No. of reflections	2657	13 281	4198	1750
No. of parameters	118	182	190	100
H-atom treatment	Constrained to parent site	Riding model	Constrained to parent site	Constrained to parent site
Weighting scheme	$w = 1/[\sigma^2(F_o^2) + (0.048P)^2P]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.098P)^2 + 8.436P]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.0505P)^2]$, where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.0725P)^2]$, where $P = (F_o^2 + 2F_c^2)/3$
(Δ/σ) _{max}	0.001	0.001	0.001	0.001
$\Delta\rho_{\max}$, $\Delta\rho_{\min}$ (e Å ⁻³)	1.26, –1.31	3.35, –2.06	0.59, –0.92	0.45, –0.16

Computer programs used: APEX2 (Bruker, 2004), SMART Version 5.625 (Bruker, 2001a), SMART Version 5.624 (Bruker, 2001a), COLLECT (Hooft, 1998), SAINT Version 6.36a (Bruker, 2000), DIRAX (Duisenberg, 1992), SHELXTL (Bruker, 2001b), HKL (Otwinowski & Minor, 1997), SHELXS97 (Sheldrick, 1990), SIR92 (Altomare *et al.*, 1994), SHELXL97 (Sheldrick, 1998), PLATON (Spek, 2003), enCIFer (Allen *et al.*, 2004).

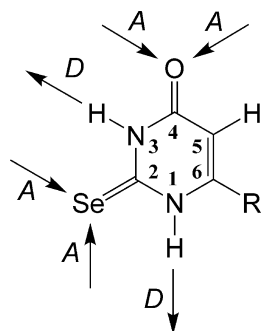
units of (1) and (11) each comprise single molecules, those of (2), (3), (4) and (4·CH₂Cl₂) comprise three, two, three and two molecules, respectively. The molecular structures of the alkylselenouracils and of 6-*n*-propyluracil are unremarkable, and the molecular structure of (4) is not affected by recrystallization as (4·CH₂Cl₂) from dichloromethane. The C–Se bond distances in (1)–(4) and in (4·CH₂Cl₂) vary from 1.824 (2) to 1.848 (6) Å (Table 2). The two C=O bond distances found in (11) are almost equal [1.2286 (16), 1.2387 (16) Å; Table 4] and are within the range of the C=O bonds found for other similar compounds such as 1,3-di-

methyl-6-*R*-trisubstituted uracils [*R* = Me, 1.233 (3) Å; *R* = Et, 1.224 (2) Å; *R* = ⁿPr, 1.212 (3), 1.221 (3), 1.220 (3) Å; *R* = ⁿBu, 1.220 (5), 1.226 (5) Å; Suwinska, 1995] and triethylammonium 2,4-dioxo-6-(1,1,2,2,3,3-hexafluoropropyl)-5-benzylsulfonyl-3*H*-2,4-dihydropyrimidine [1.235 (3), 1.237 (2) Å] (Timoshenko *et al.*, 2002).

3.2. Extended structures of (1)–(4), (4·CH₂Cl₂) and (11)

The extended structures formed by the alkylselenouracils are worthy of comment; they depend primarily on hydrogen-

bonding interactions between the two N–H donor sites and the oxygen and selenium acceptor sites of the heteroatomic ring [see Scheme (II)]. Pertinent distances and angles associated with the hydrogen-bonding interactions in (1)–(4), (4·CH₂Cl₂) and (11) are included in Table 3.



Analysis of the hydrogen-bonding parameters in these compounds confirms a range of intermolecular contacts with similar numbers of N1–H···O and N3–H···O and of N1–H···Se and N3–H···Se interactions.

The extended structure of (1) is a two-dimensional sheet of topology 6³ (Fig. 7). A centrosymmetrically related pair of N3–H···Se contacts form an $R_2^2(8)$ ring which links two molecules, thereby forming a dimeric unit. Each dimeric unit is bridged to four adjacent dimeric units through four single N1–H···O contacts to give a sheet of six-membered rings, each of which involves an $R_6^0(28)$ hydrogen-bonded motif (Table 3). The 6³ topology of the sheet, which is aligned

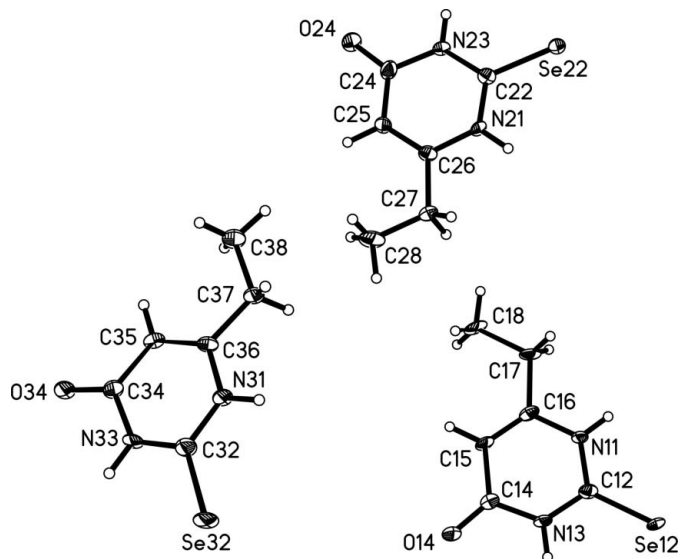


Figure 2
Displacement ellipsoid plot of (2) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level. There are three independent molecules in the asymmetric unit.

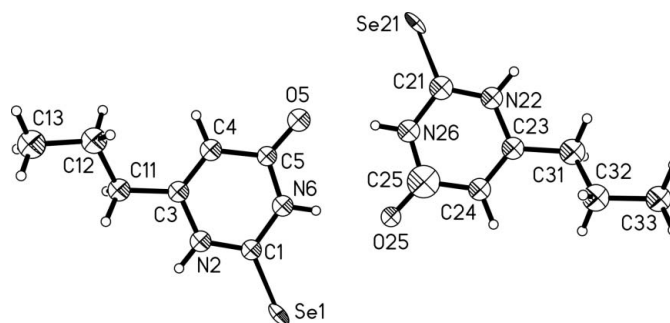


Figure 3
Displacement ellipsoid plot of (3) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level. There are two independent molecules in the asymmetric unit.

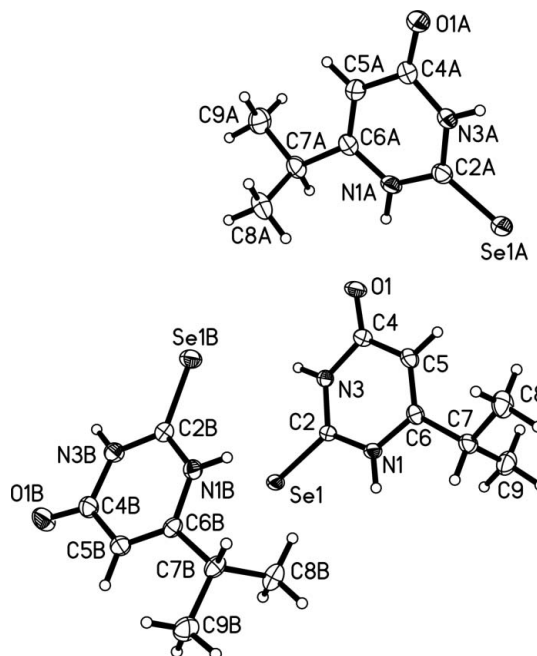


Figure 4
Displacement ellipsoid plot of (4) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level. There are three independent molecules in the asymmetric unit.

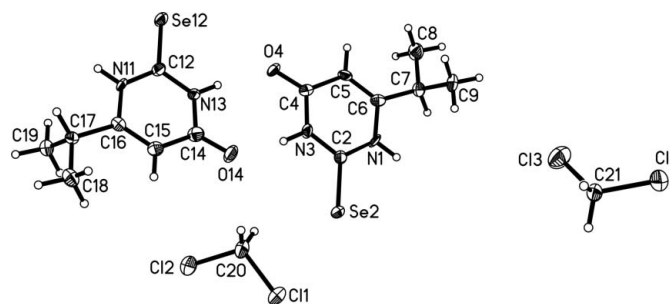


Figure 5
Displacement ellipsoid plot of (4·CH₂Cl₂) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level. Two independent molecules of (4) are solvated by two dichloromethane molecules in the asymmetric unit.

Table 2
Molecular geometry parameters (Å, °) for (1) (2), (3), (4) and (4·CH₂Cl₂).

(1)	(2)	(3)	(4)	(4·CH ₂ Cl ₂)					
Se2—C2	1.832 (2)	Se12—C12	1.835 (6)	Se1—C1	1.837 (16)	Se1—C2	1.843 (2)	Se2—C2	1.829 (6)
N1—C2	1.352 (3)	Se22—C22	1.848 (6)	Se21—C21	1.835 (16)	Se1A—C2A	1.824 (2)	Se12—C12	1.825 (5)
N3—C2	1.353 (3)	Se32—C32	1.828 (7)	N2—C1	1.32 (2)	Se1B—C2B	1.835 (2)	N1—C2	1.345 (7)
O4—C4	1.233 (3)	N11—C12	1.354 (8)	N2—C3	1.40 (2)	N1—C2	1.348 (3)	N3—C2	1.351 (7)
		N21—C22	1.341 (8)	N22—C2	1.39 (2)	N1A—C2A	1.354 (3)	N11—C12	1.355 (7)
Se2—C2—N3	121.99 (16)	N31—C32	1.352 (9)	N6—C1	1.39 (2)	N1B—C2B	1.342 (3)	N13—C12	1.349 (7)
Se2—C2—N1	121.90 (15)	N13—C12	1.352 (7)	N26—C21	1.39 (2)	N1—C6	1.389 (3)	O4—C4	1.232 (7)
		N23—C22	1.349 (7)	O5—C5	1.23 (2)	N1A—C6A	1.383 (3)	O14—C14	1.233 (7)
		N33—C32	1.349 (8)	O25—C25	1.24 (2)	N1B—C6B	1.389 (3)		
		O14—C14	1.240 (8)	Se1—C1—N2	123.3 (13)	C2—N3	1.341 (3)	Se2—C2—N1	123.0 (4)
		O24—C24	1.242 (8)	Se1—C1—N6	121.7 (12)	C2A—N3A	1.351 (3)	Se2—C2—N3	121.5 (4)
		O34—C34	1.227 (7)	Se21—C21—N22	123.9 (13)	C2B—N3B	1.353 (3)	Se12—C12—N11	123.5 (4)
				Se21—C21—N26	121.2 (12)	N3—C4	1.400 (3)	Se12—C12—N13	121.1 (4)
		Se12—C12—N11	123.2 (4)			N3A—C4A	1.391 (3)		
		Se22—C22—N21	121.0 (4)			N3B—C4B	1.395 (3)		
		Se32—C32—N31	123.7 (5)			O1—C4	1.230 (3)		
		Se12—C12—N13	121.6 (4)			O1A—C4A	1.237 (3)		
		Se22—C22—N23	122.8 (4)			O1B—C4B	1.226 (3)		
		Se32—C32—N33	121.3 (5)						
						Se1—C2—N1	120.70 (17)		
						Se1A—C2A—N1A	123.65 (18)		
						Se1B—C2B—N1B	122.91 (17)		
						Se1—C2—N3	122.59 (17)		
						Se1A—C2A—N3A	121.33 (18)		
						Se1B—C2B—N3B	121.19 (17)		

parallel to the (10 $\bar{2}$) plane, results in a brick-wall architecture (Fig. 7).

The extended structures of (2) and (4) are analogous. They crystallize in the same space group with similar cell dimensions, the only difference being the alkyl group. Their extended structure (Fig. 8) is based on a chain of eight-membered rings which is aligned in the [101] direction in the (21 $\bar{2}$) plane. The three molecules in the asymmetric unit are involved in different hydrogen-bonding motifs; the first (numbered C11 *etc.*) acts as a two-donor one-acceptor (2D–1A) species; the second (numbered C21 *etc.*) acts as a two-donor three-acceptor species (2D–3A); the third (numbered C31 *etc.*) acts as a two-donor two-acceptor species (2D–2A). Each eight-membered ring is centrosymmetric and comprises two 2D–1A molecules, two 2D–2A molecules and four 2D–3A molecules. Pairs of centrosymmetrically related 2D–3A molecules are common to two adjacent rings, thereby giving

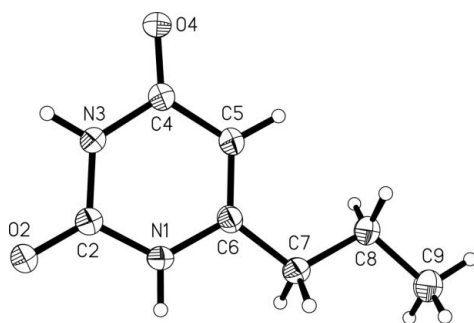


Figure 6
Displacement ellipsoid plot of (1) showing the atomic-numbering scheme, with ellipsoids drawn at the 50% probability level.

the correct stoichiometry. The hydrogen-bonding interactions leading to the eight-membered ring comprise an $R_8^6(32)$ motif (Table 3). The aliphatic residues of the two 2D–1A and two 2D–2A molecules are located on the edges of the chain, preventing inter-chain hydrogen-bond formation (Fig. 8).

In (3), $R_2^2(8)$ hydrogen-bonding motifs link the two molecules of the asymmetric unit, which alternate along the *b* axis to form a one-dimensional chain (Fig. 9). The $R_2^2(8)$ motifs differ in that one has oxygen acceptors and the other has selenium acceptors. The 1:1 adduct (4·CH₂Cl₂) has a similar

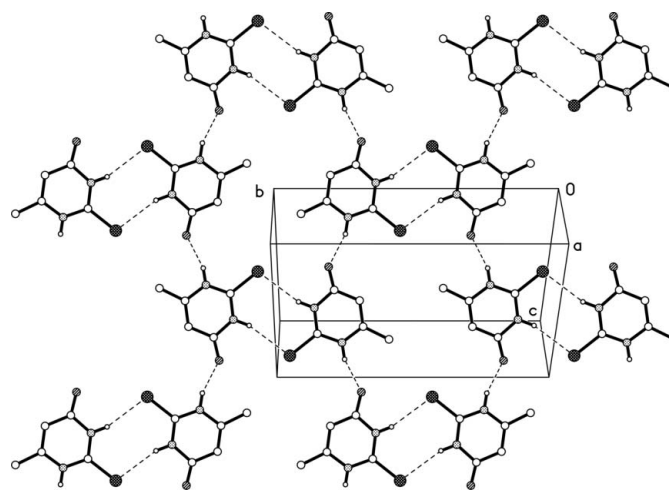


Figure 7
View of part of a sheet in the crystal structure of (1) showing the intermolecular N–H...O and pairwise N–H...Se interactions. Atoms are identified as follows: Se, large cross-hatched circles; O, hatched circles; N, dotted circles; C, intermediate open circles; H, small open circles.

Table 3

Hydrogen-bonding parameters (Å, °) for (1)–(4), (4-CH₂Cl₂) and (11) [N–H = 0.88 Å; C–H = 0.95 Å; λ = 0.6775 Å for (1), (3) and (10)].

Interaction	H...Y (Å)	X...Y (Å)	X–H...Y (°)
(1)			
N1–H1A...O4	1.91	2.761 (2)	162
N3–H3A...Se2	2.63	3.4825 (14)	162
(2)			
N11–H11...O34	1.93	2.805 (6)	171
N23–H23...O14	1.99	2.868 (7)	172
N31–H31...O24	1.97	2.829 (7)	166
N13–H13...Se22	2.47	3.340 (5)	169
N21–H21...Se22	2.58	3.442 (5)	167
N33–H33...Se12	2.55	3.417 (6)	170
(3)			
N6–H6N...O25	1.96	2.80 (2)	161
N26–H26N...O5	1.99	2.84 (2)	162
N2–H2N...Se21	2.60	3.44 (2)	160
N22–H22N...Se1	2.59	3.46 (2)	167
(4)			
N1A–H1AA...O1	2.00	2.860 (3)	167
N1B–H1BA...O1A	1.92	2.797 (3)	177
N3–H3A...O1B	1.91	2.782 (3)	169
N1–H1A...Se1	2.60	3.460 (2)	165
N3A–H3AA...Se1B	2.60	3.430 (3)	157
N3B–H3BA...Se1	2.56	3.428 (2)	167
(4-CH₂Cl₂)			
N3–H3...O14	1.97	2.831 (6)	167
N13–H13...O4	1.98	2.854 (6)	172
N1–H1...Se12	2.55	3.424 (5)	175
N11–H11...Se2	2.57	3.448 (5)	176
C21–H21A...O4	2.34	3.230 (7)	150
(11)			
N1–H1A...O4	1.91	2.771 (2)	166
N3–H3A...O2	2.01	2.856 (2)	161
C9–H9A...O4	2.52	3.491 (2)	169

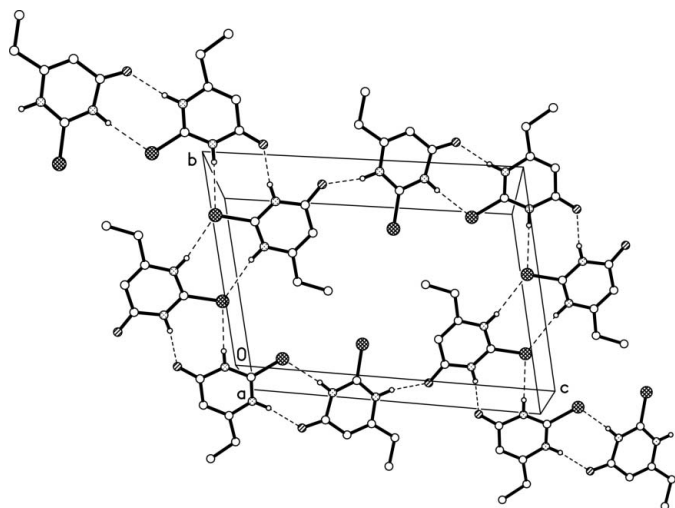


Figure 8

View of part of a sheet in the crystal structure of (2) showing the intermolecular N–H...O and N–H...Se interactions. Atoms are identified as in Fig. 7.

extended structure (Fig. 10) to that of (3). In this case, however, the chain is aligned along the *b* axis. One of the dichloromethane molecules is locked in position by a relatively short C–H...O contact (Table 3).

It is interesting to note that a different extended structure is adopted by 6-*n*-propyluracil (11), the uracil corresponding to (3). The asymmetric unit of (11) comprises a single molecule (Fig. 6), six of which combine to form a six-membered ring (Fig. 11). Linking each pair of molecules is an $R_2^2(8)$ hydrogen-bonding motif, which forms part of the $R_6^6(24)$ hydrogen-bonding motif (Table 3), generating the inner diameter of the six-membered ring. The rings assemble on a sheet parallel to the (001) plane. The only inter-ring interactions are a centrosymmetric pair of relatively long C–H...O contacts between the pendant CH₃ groups and carbonyl O atoms (Table 3). Six-membered ring formation is possible owing to the fact that the $R_2^2(8)$ motifs in (11) [see Scheme (III*a*)], which comprise two N–H...O contacts, are symmetrical and generate an internal ring angle of 120°. If (3) were to adopt a

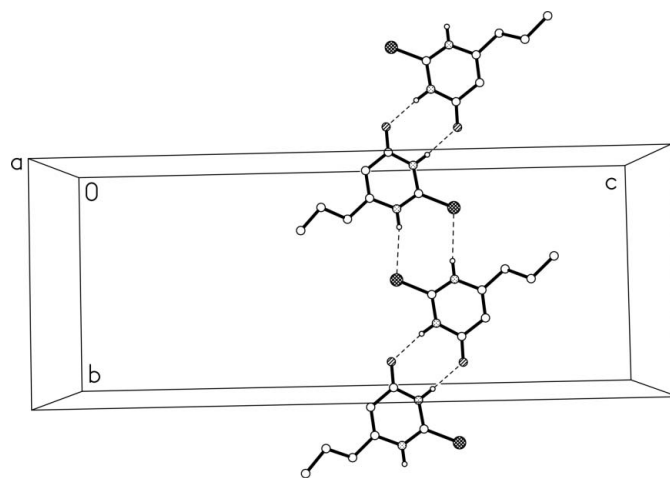


Figure 9

View of part of a chain in the crystal structure of (3) showing the intermolecular N–H...O and N–H...Se interactions. Atoms are identified as in Fig. 7.

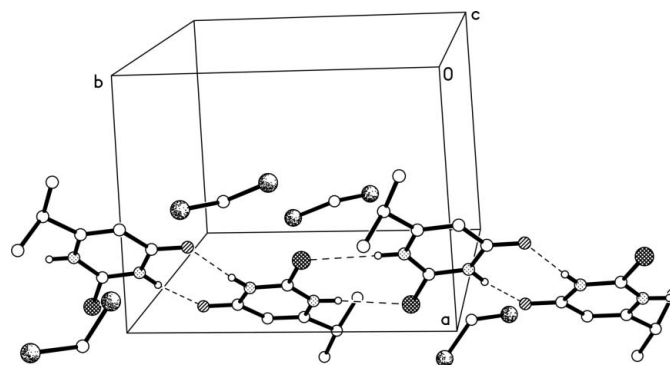


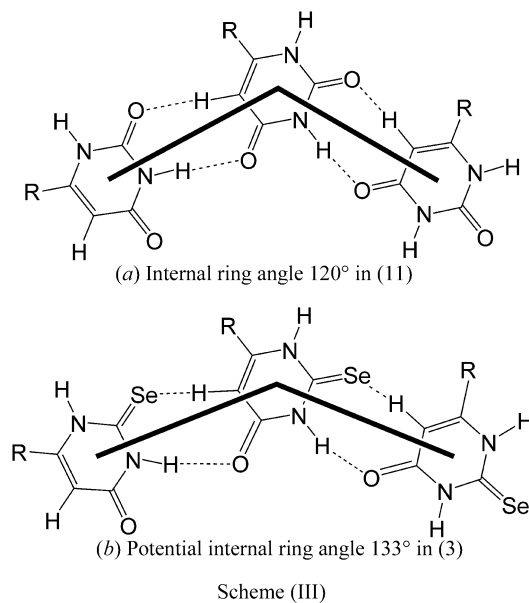
Figure 10

View of part of a chain in the crystal structure of (4-CH₂Cl₂) showing the intermolecular N–H...O and N–H...Se interactions. Atoms are identified as in Fig. 7.

Table 4
Molecular geometry parameters (Å, °) for (7), (9·2H₂O), (10) and (11).

(7)	(9·2H ₂ O)	(10)	(11)				
I1—I2	2.8928 (10)	Se2—Se2	2.4328 (9)	Se2—Se2	2.4427 (6)	O2—C2	1.2286 (16)
I1—Se	2.7807 (11)	Se2—C2	1.925 (4)	Se2—C2	1.922 (4)	O4—C4	1.2387 (16)
Se—C2	1.876 (6)	N1—C2	1.283 (6)	N1—C2	1.296 (4)	N1—C2	1.3668 (15)
N1—C2	1.331 (8)	N3—C2	1.412 (6)	N3—C2	1.407 (4)	N3—C2	1.3653 (14)
N3—C2	1.334 (8)	N3—C4	1.431 (6)	N3—C4	1.421 (3)	N3—C4	1.3895 (12)
N3—C4	1.405 (8)	O4—C4	1.233 (5)	O4—C4	1.236 (4)		
O4—C4	1.213 (8)	N3—C12	1.420 (5)	N3—C12	1.436 (4)	O2—C2—N3	122.39 (11)
		N11—C12	1.296 (4)	N11—C12	1.271 (4)	O2—C2—N1	122.45 (11)
I2—I1—Se	176.75 (2)	N13—C12	1.337 (3)	N13—C12	1.356 (4)	O4—C4—N3	119.56 (11)
I1—Se—C2	96.9 (2)	N13—C14	1.400 (3)	N13—C14	1.396 (4)	O4—C4—C5	125.10 (9)
		O14—C14	1.232 (3)	O14—C14	1.230 (4)	N1—C2—N3	115.15 (8)
		Se2—Se2—C2	88.99 (14)	Se2—Se2—C2	89.44 (8)		
		Se2—C2—N1	114.7 (3)	Se2—C2—N1	122.2 (2)		
		Se2—C2—N3	121.6 (3)	Se2—C2—N3	115.1 (2)		
		C2—Se2—Se2—C2	−179.98 (18)	C2—Se2—Se2—C2	180		

similar structure, the corresponding $R_2^2(8)$ motifs would be unsymmetrical as they would comprise one short N—H···O contact and one long N—H···Se contact. Analysis of the structure of (2), which contains such assemblies, gives a value of 133° for the internal ring angle [see Scheme (IIIb)], which does not permit ring formation as it falls between the values required for seven- and eight-membered rings, namely 128.6 and 135°, respectively. Hence, (3) adopts a one-dimensional chain architecture with alternating $R_2^2(8)$ rings with pairs of oxygen and selenium acceptors (Fig. 9).



3.3. Molecular structures of (7), (9·2H₂O) and (10)

Displacement ellipsoid plots of compounds (7), (9·2H₂O) and (10) are shown in Figs. 12, 13 and 14, while their selected bond lengths and angles are listed in Table 4. Compound (7) exhibits the so-called ‘spoke’ structure typical of iodine adducts of selenium-containing compounds. The I—Se—C—N torsion angles [−0.5 (6)° and −179.3 (6)° for (7)], together

with the C—Se—I and Se—I—I angles (Table 4), are consistent with a planar arrangement. This planarity, which is typical of charge-transfer complexes with a ‘spoke’ structure, is

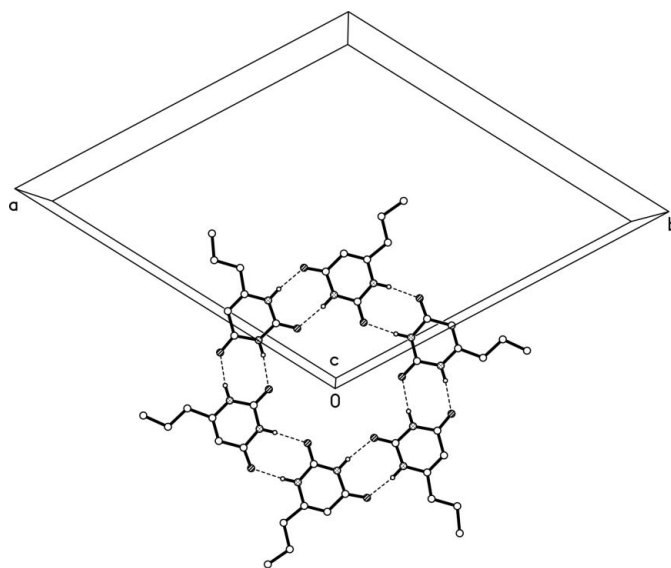


Figure 11
View of part of a sheet of (11) showing the intermolecular N—H···O interactions. Atoms are identified as in Fig. 7.

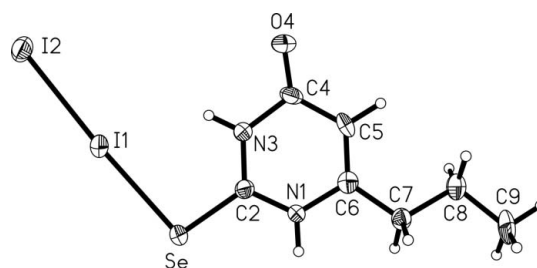


Figure 12
Displacement ellipsoid plot of (7) showing the atom-numbering scheme, with ellipsoids drawn at the 50% probability level.

Table 5

Hydrogen-bonding parameters (Å, °) for (7), (9·2H₂O) and (10) (N—H = 0.88 Å; C—H = 0.95 Å).

Interaction	H···Y (Å)	X···Y (Å)	X—H···Y (°)
(7)			
N1—H1N···O4	1.85	2.698 (8)	161
N3—H3N···I1	2.66	3.371 (5)	139
(9·2H ₂ O)			
N13—H13A···O4	1.88	2.560 (5)	133
N13—H13A···O4'	2.36	3.166 (5)	152
C5—H5A···O14	2.48	3.304 (5)	145
O1···O14		2.861 (5)	
O1···O1'		2.741 (5)	
O1···O1'		2.770 (5)	
(10)			
N13—H13A···O4	1.84	2.534 (3)	134
C5—H5A···O4	2.59	3.535 (4)	176
C15—H15A···O14	2.61	3.562 (4)	175

enhanced by an intramolecular N—H···I hydrogen bond between N3 and I1 (Table 5; Fig. 15).

The I—I interatomic distance of 2.8928 (10) Å in (7) is longer than that in either the gas phase (2.677 Å; Pauling, 1960) or crystalline diiodine [2.715 (6) Å at 110 K; van Bolhuis *et al.*, 1967], presumably owing to the Se···I interaction. It is, however, the shortest such distance measured for a diiodine–selenoamide complex (Antoniadis *et al.*, 2006) suggesting minimal perturbation resulting from the long Se···I contact. Interestingly, correlation of the available Se···I and I···I distance data (see Antoniadis *et al.*, 2006) shows there to be a linear relationship between the two, which is quantified by the expression: $d(\text{Se}\cdots\text{I}) = -0.7981 d(\text{I}\cdots\text{I}) + 5.0983$, $R^2 = 0.9805$. The I—I bond order of 0.547 calculated for (7) using the expression of Pauling (1960) is the highest such bond order for

selenoamide–diiodine complexes. All these data are consistent with a very weak Se···I interaction. Bigoli *et al.* (1996, 1999) have separated iodine adducts of sulfur donors into three classes on the basis of I···I bond order. Applying the same criteria to iodine adducts of selenium donors, (7) falls into the

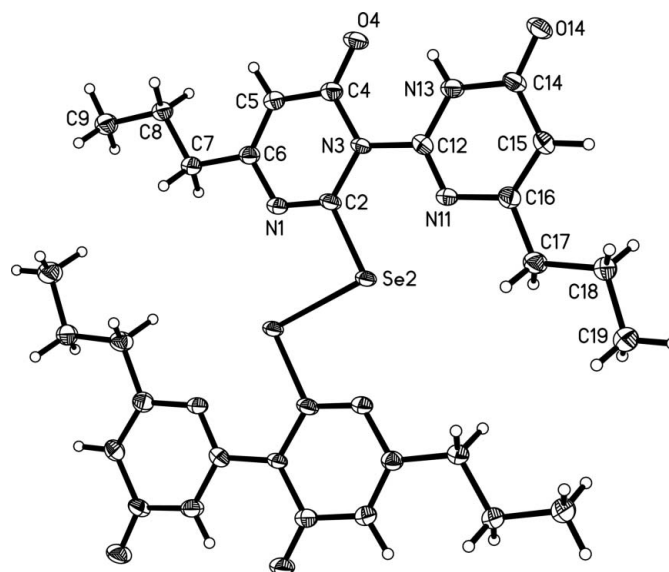


Figure 14
Displacement ellipsoid plot of molecule (10) showing the atomic-numbering scheme, with ellipsoids drawn at the 50% probability level. The diselenide link is formed between Se2 and its symmetry equivalent at $(-x, -y, -z)$.

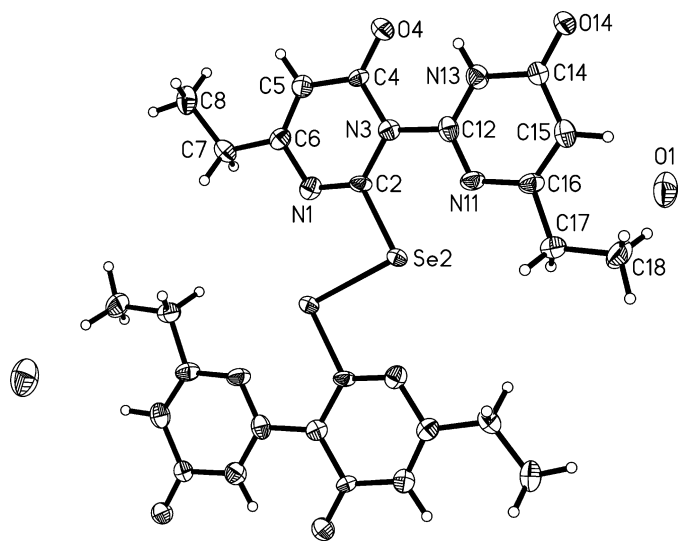


Figure 13
Displacement ellipsoid plot of molecule (9·2H₂O) showing the atomic-numbering scheme, with ellipsoids drawn at the 50% probability level. The diselenide link is formed between Se2 and its symmetry equivalent at $(1-x, -y, 1-z)$. The water H atoms were not located.

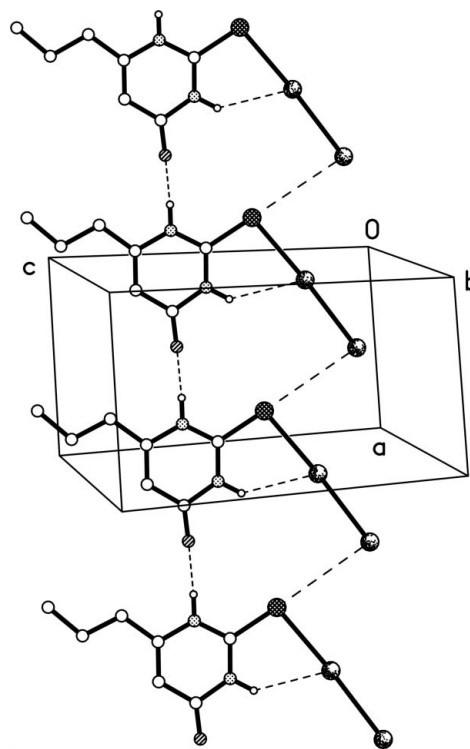


Figure 15
View of part of a chain in the crystal structure of (7) showing the intermolecular N—H···O and C—H···I and Se···I interactions. Atoms are identified as in Fig. 7.

intermediate, rather than the weakest, classification which would require an $I \cdots I$ distance of 3.92 Å, which is close to double the van der Waals radius of iodine (1.98 Å). Therefore, a third type of Se–I adduct seems very unlikely. Although the C–Se distance in (7) [1.876 (6) Å] is slightly longer than that for the corresponding free selenouracil (3) [average 1.836 (11) Å; Table 2], it is similar to the C–Se distances found for the other charge-transfer complexes (see Antoniadis *et al.*, 2006).

Re-crystallization of (6) and (7) from acetone solutions results in the formation of the diselenides (9·2H₂O) and (10), respectively [see Scheme (I)]. Their molecular structures consist of two centrosymmetrically related selenoamide ligands, which have been *N*-substituted by a deselenated ligand molecule, linked through an Se–Se bond to form the diselenide (Figs. 13 and 14). The presence of a crystallographic inversion centre between the selenoamides fixes the C–Se–Se–C torsion angle at 180°, which is unusual for diselenoamides. The majority of those listed have torsion angles between 50 and 93°, the only other example of a 180° angle

being that in *N,N,N',N'*-tetraethylthiuramdiseleide (Dietzsch *et al.*, 1998).

The Se–Se bond distances in (9·2H₂O) [2.4328 (9) Å] and (10) [2.4427 (6) Å] fall within the range observed in other diselenoamides (2.34–2.59 Å; Antoniadis *et al.*, 2006). The C–Se bond lengths in (9·2H₂O) [1.925 (4) Å] and (10) [1.922 (4) Å], although longer than those of the corresponding free ligands (2) [mean C–Se 1.838 (5) Å] and (3) [mean C–Se 1.836 (11) Å], fall in the range of other such compounds (1.874–1.952 Å; Antoniadis *et al.*, 2006). Molecules (9·2H₂O) and (10) are neutral diselenoamides with two unequal C–N bond distances, as expected for such compounds. These bond distances become equal in the case of ionic diselenoamides. The Se–Se–C bond angles in (9·2H₂O) [88.99 (14)°] and (10) [89.44 (8)°] are at the lower end of the range found for diselenoamides.

3.4. Extended structures of (7), (9·2H₂O) and (10)

The extended structure of (7) is based on a linear chain aligned along the *b* axis (Fig. 15). The principal contact between adduct molecules is an intermolecular N–H···O hydrogen bond (Table 5). It is supported by an interaction between the terminal iodine and the selenium atom [Se···I2 ($x - 1, y, z$) 3.862 (2) Å; I1–I2···Se ($1 + x, y, z$) 87.02 (3); C2–Se···I2 ($x - 1, y, z$) 171.3 (2); I1–Se···I2 ($x - 1, y, z$) 83.82 (3); C2–Se···I1 96.9 (2)°].

The extended structures of (9·2H₂O) and (10) are similar. They form chains aligned along the [210] and [101] directions, respectively, as shown in Figs. 16 and 17, respectively. The principal links between molecules in (9·2H₂O) are N–H···O and in (10) are C–H···O hydrogen bonds. The reason for the

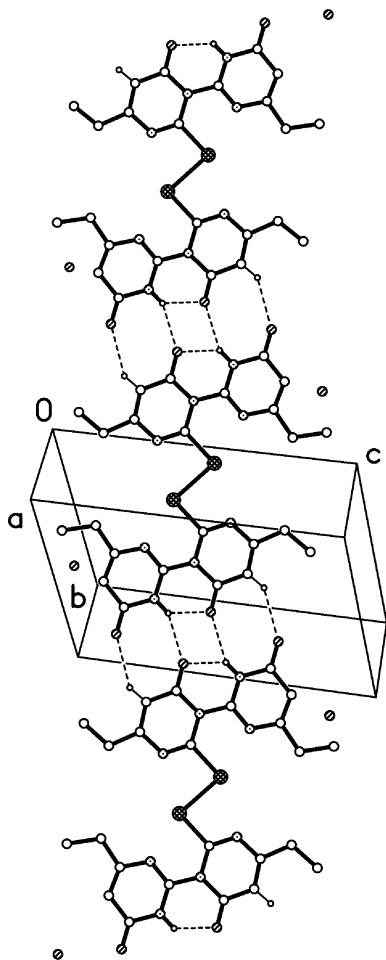


Figure 16
View of part of a sheet of (9·2H₂O) showing the inter- and intramolecular N–H···O and intermolecular C–H···O interactions. Atoms are identified as in Fig. 7.

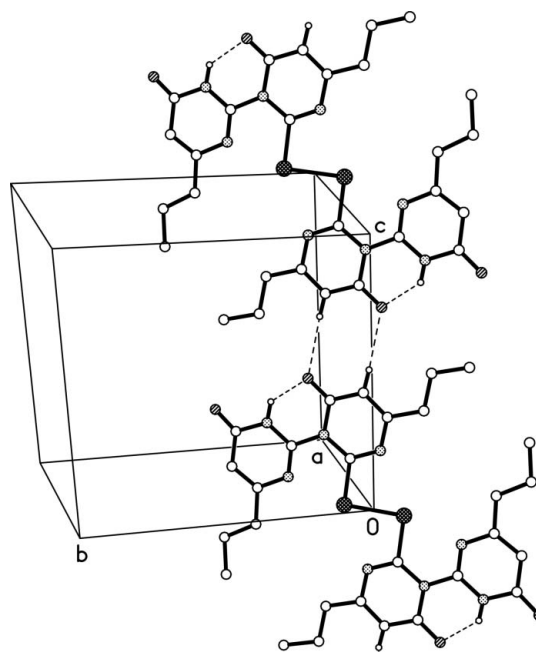


Figure 17
View of part of a sheet of (10) showing the intramolecular N–H···O and intermolecular C–H···O interactions. Atoms are identified as in Fig. 7.

difference is not clear. In the case of (9·2H₂O), the N—H···O contact is quite long at 3.166 (5) Å, and by inference weak, owing to bifurcation of the hydrogen bond, the other contact being an intramolecular interaction of 2.560 (5) Å. The intermolecular contact is supported by a relatively short C—H···O hydrogen bond [C···O 3.304 (5) Å]. In the case of (10), the N—H···O intramolecular contact [N···O 2.534 (3) Å] is an independent interaction, the molecules being linked by a pair of centrosymmetrically related, somewhat long, C—H···O hydrogen bonds [C···O 3.535 (4) Å] which form an $R_2^2(8)$ motif (Fig. 17).

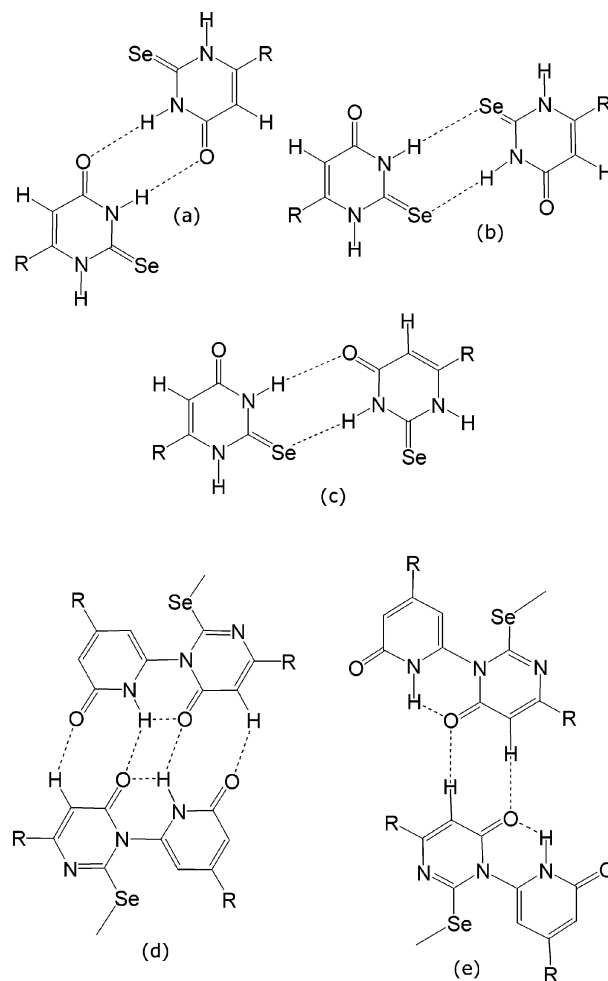
In both structures the chains are stacked to give two-dimensional sheet architectures parallel to ($\bar{1}20$) for (9·2H₂O) and to ($\bar{1}21$) for (10). In the case of (9·2H₂O) the chains are linked through hydrogen-bonded water molecules which not only form the sheet but also link the sheets into a three-dimensional structure. Although it was not possible to locate the water H atoms, the proposed interactions are supported by the geometries around the O atoms. Thus, O···O distances range from 2.741 (5) through 2.770 (5) to 2.861 (5) Å, while O···O···O angles range from 109.9 (2) through 122.5 (3) to 123.7 (2)°. There are no aromatic stacking interactions owing to a staggered molecular arrangement. The only other possible interaction linking the sheets is an Se···Se contact, but the interatomic distance is extremely long [4.125 (2) Å]. In the case of (10) the chains are linked into two-dimensional sheets by relatively long C—H···O contacts [C···O 3.562 (4) Å; Table 5]. Hydrogen-bonded contacts between sheets are not feasible and the staggered molecular arrangement precludes aromatic stacking interactions. The only possible interaction is an Se···Se contact, but the interatomic distance of 3.961 (2) Å is rather long.

4. Concluding remarks

The architectures of the extended structures of the 6-alkyl-2-selenouracils (1)–(4) and (4·CH₂Cl₂) rely heavily on $R_2^2(8)$ hydrogen-bonded rings. All three types of rings, namely those based on two oxygen acceptors [O/O; see Scheme (IVa)], two selenium acceptors [Se/Se; see Scheme (IVb)] and a mixture of one oxygen and one selenium acceptor [O/Se; see Scheme (IVc)], are found in the 6-alkyl-2-selenouracils. The most common is the Se/Se ring; it appears in the extended structures of all five compounds [(1)–(4) and (4·CH₂Cl₂)]. The other two types of ring appear twice only, the O/O ring in (3) and (4·CH₂Cl₂) and the O/Se ring in (2) and (4). The simplest structures are those of (3) and (4·CH₂Cl₂), where alternating Se/Se and O/O rings give rise to one-dimensional chain constructs (Figs. 9 and 10). In (1), Se/Se rings form dimers which are linked through N—H···O hydrogen bonds to give a two-dimensional sheet of topology δ^3 (Fig. 7). In (2) and (4), Se/Se and O/Se rings alternate to form chains of 6-alkyl-2-selenouracils, which are linked through N—H···O hydrogen bonds to give a two-dimensional sheet (Fig. 8). O/O $R_2^2(8)$ hydrogen-bonded rings are also found in the extended structure of 6-*n*-propyl-2-uracil (11), where they generate a six-

membered paddle-wheel ring with pendant propyl groups (Fig. 11).

The extended structures of the diselenide oxidation products (9·2H₂O and 10) also depend on O/O $R_2^2(8)$ hydrogen-bonded rings. However, in (9·2H₂O) the donors are a mixture of C—H and N—H moieties [see (IVd)] while in (10) they are exclusively C—H donors. In both cases the molecules are linked to give rise to chain architectures (Figs. 16 and 17).



Scheme (IV)

The O atoms of all five 6-alkyl-2-selenouracil and of the single 6-*n*-propyl-2-uracil structures act as acceptors to just one hydrogen-bonded contact. The Se atoms behave similarly except in (2) and (4). Of the Se atoms of the three molecules in the asymmetric unit of (2) and (4), one acts as a dual acceptor, one as a single acceptor while the third is not involved in any hydrogen bonds (Fig. 8). The situation in the oxidation products (9·2H₂O and 10) is complicated by the existence of intramolecular N—H···O hydrogen bonds. In (9·2H₂O) one oxygen acts as a dual acceptor to intra- and intermolecular N—H···O hydrogen bonds and a second oxygen acts as a single acceptor in an intermolecular C—H···O hydrogen bond. All three contacts generate the one-dimensional chain structure (Fig. 16). In (10), one oxygen acts as a dual acceptor

to an intramolecular N—H···O and an intermolecular C—H···O hydrogen bond to generate the one-dimensional chain structure, while a second oxygen acts as a single acceptor in an interchain C—H···O hydrogen bond (Fig. 17).

We thank the Marie Curie Foundation for funding (to CDA) as part of the Marie Curie Host Fellowship HPMT-CT-2001-00376 *COSMIC*, and the graduate program in Bioinorganic Chemistry based at Ioannina, Greece and coordinated by NH. We thank EPSRC (UK) for the award of diffractometers and for data collection by the EPSRC National Service for Crystallography at the University of Southampton, UK, and at the Daresbury Laboratory Synchrotron Radiation Source. MS acknowledges receipt of a Royal Society–Leverhulme Trust Senior Research Fellowship and of a Royal Society Wolfson Merit Award.

References

- Allen, F. H., Johnson, O., Shields, G. P., Smith, B. R. & Towler, M. (2004). *J. Appl. Cryst.* **37**, 335–338.
- Altomare, A., Burla, M. C., Camalli, M., Cascarano, G., Giacovazzo, C., Guagliardi, A. & Polidori, G. (1994). *J. Appl. Cryst.* **27**, 435.
- Antoniadis, C. D., Hadjikakou, S. K., Hadjiliadis, N., Baril, M. & Butler, I. S. (2006). *Chem. Eur. J.* DOI: 10.1002/chem.200501455.
- Aragoni, M. C., Arca, A., Demartin, F., Devillanova, F. A., Garau, A., Isaia, F., Lippolis, V. & Verani, G. (1999). *Trends Inorg. Chem.* **6**, 1–18.
- Aragoni, M. C., Arca, M., Blake, A. J., Devillanova, F. A., Du Mont, W.-W., Garau, A., Isaia, F., Lippolis, V., Verani, G. & Wilson, C. (2001). *Angew. Chem. Int. Ed.* **40**, 4229–4232.
- Bent, H. A. (1968). *Chem. Rev.* **68**, 587–648.
- Bigoli, F., Deplano, P., Ienco, A., Mealli, C., Mercuri, M. L., Pellinghelli, M. A., Pintus, G., Saba, G. & Trogu, E. F. (1999). *Inorg. Chem.* **38**, 4626–4636.
- Bigoli, F., Deplano, P., Mercuri, M. L., Pellinghelli, M. A., Sabatini, A., Trogu, E. F. & Vacca, A. (1996). *J. Chem. Soc. Dalton Trans.* pp. 3583–3589.
- Blake, A. J., Devillanova, F. A., Garau, A., Gilby, L. M., Gould, R. O., Isaia, F., Lippolis, V., Parsons, S., Radek, C. & Schröder, M. (1998). *J. Chem. Soc. Dalton Trans.* pp. 2037–2046.
- Blake, A. J., Devillanova, F. A., Gould, R. O., Li, W.-S., Lippolis, V., Parsons, S., Radek, C. & Schröder, M. (1998). *Chem. Soc. Rev.* **27**, 195–205.
- Blake, A. J., Gould, R. O., Parsons, S., Radek, C. & Schröder, M. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 2374–2376.
- Blake, A. J., Li, W.-S., Lippolis, V., Parsons, S., Radek, C. & Schröder, M. (1998). *Angew. Chem. Int. Ed. Engl.* **37**, 293–296.
- Bjorvatten, T. (1963). *Acta Chem. Scand.* **17**, 2292–2300.
- Bolhuis, F. van, Koster, P. B. & Mighelsen, T. (1967). *Acta Cryst.* **23**, 90–91.
- Bruker (2000). *SAINT*. Version 6.36a. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2001a). *SMART*. Version 5.624 or 5.625. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2001b). *SHELXTL*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2004). *APEX2*. Version 1.0–27. Bruker AXS Inc., Madison, Wisconsin, USA.
- Chao, G. Y. & McCullough, J. D. (1961). *Acta Cryst.* **14**, 940–945.
- Dahl, T. & Hassel, O. (1965). *Acta Chem. Scand.* **19**, 2000–2001.
- Dasent, W. E. (1965). *Non-Existent Compounds*, p. 162. New York: Marcel Dekker.
- Dietzsch, W., Sieler, J., Meiler, W. & Robien, W. (1998). *Phosphorus, Sulfur Silicon Relat. Elem.* **38**, 293.
- Duisenberg, A. J. M. (1992). *J. Appl. Cryst.* **25**, 92–96.
- Holmesland, O. & Römning, C. (1966). *Acta Chem. Scand.* **20**, 2601.
- Hooft, R. W. W. (1998). *Collect. Nonius BV*, Delft, The Netherlands.
- Hope, H. & McCullough, J. D. (1962). *Acta Cryst.* **15**, 806–807.
- Maddox, H. & McCullough, J. D. (1966). *Inorg. Chem.* **5**, 522–526.
- Mont, W.-W. du, von Salzen, A. M., Ruthe, F., Seppala, E., Mugesh, G., Devillanova, F. A., Lippolis, V. & Kuhn, N. (2001). *J. Organomet. Chem.* **623**, 14–28.
- Otwinowski, Z. & Minor, W. (1997). *Methods Enzymol.* **276**, 307–326.
- Pauling, L. (1960). *The Nature of the Chemical Bond*, 3rd ed. Ithaca, NY: Cornell University Press.
- Sheldrick, G. M. (1990). *Acta Cryst.* **A46**, 467–473.
- Sheldrick, G. M. (1998). *SHELXL97–2*. University of Göttingen, Göttingen, Germany.
- Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.
- Suwinska, K. (1995). *Acta Cryst.* **B51**, 248–254.
- Timoshenko, V. M., Nikolin, Y. V., Chernega, A. N. & Shermolovich, Y. G. (2002). *Eur. J. Org. Chem.* pp. 1619–1627.